

STATE LIBRARY OF QUEENSLAND

Year 9 Chemical Sciences

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Overview

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Before you get started

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Once you have joined State Library, <u>log in</u> to your account in One Search so you can use the links in this research guide to access the featured collection items.

You can search our <u>OneSearch catalogue</u> or begin exploring by clicking on some of the featured items contained in this research guide.

E-Books











The Atomic Model

Explain how the model of the atom changed following the discovery of electrons, protons and neutrons and describe how natural radioactive decay results in stable atoms.

- Comparing the mass and charge of protons, neutrons and electrons.
- Examining how the discovery of electrons, protons and neutrons resulted from experimental evidence and answered questions related to properties and behaviours of atoms .
- Explaining that differences in the number of neutrons in atoms of the same element results in isotopes and that naturally occurring isotopes of some elements are unstable.
- Describing in simple terms how different unstable isotopes decay such as radon-222 releasing an alpha particle, iodine-131 releasing a beta particle and cobalt-60 releasing gamma radiation to form stable atoms.
- Defining half-life, examining the timescales of decay of different elements such as carbon-14 and uranium-238 and simulating or using digital simulations to examine radioactive decay including half-life.
- Investigating how radiocarbon and other dating methods have been used to establish that First Peoples of Australia have been present on the Australian continent for more than 60,000 years.
- Identifying where applications of radioactivity are used in medicine and industry such as diagnosing and treating cancer and checking for faults in materials used in aircraft and spacecraft.
- Discussing how mass and energy are connected at all scales and energy conversion processes within atomic nuclei.



<u>CHEMnetBASE</u> provides a comprehensive database of chemicals and their uses, as well as chemical equations and properties.

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<u>Queensland Museum</u> provide Learning Resources website students with many activities, fact sheets, images, and videos.







<u>Oxford Reference</u> spans 25 different subject areas, bringing together 2 million digitised entries across Oxford University Press's dictionaries, companions, and encyclopedias.

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<u>JSTOR</u> provides access to more than 12 million scholarly journal articles and eBooks, and is especially good for primary sources.



<u>Gale Interactive: Science</u> provides a comprehensive view of the most-studied science subjects. Authoritative, high-quality digital content is paired with interactive 3D models.







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The Law of Conservation of Mass

Model the rearrangement of atoms in chemical reactions using a range of representations, including word and simple balanced chemical equations, and use these to demonstrate the law of conservation of mass.

- Identifying reactants and products in chemical reactions.
- Using models and representations to show the rearrangement of atoms in chemical reactions.
- Investigating chemical reactions in closed and open systems and relating data obtained to the law of conservation of mass.
- Writing symbolic equations that are easy to balance and explaining, using the law of conservation of mass, and atoms, the rationale for balancing chemical equations.
- Investigating why most elements are not found in their elemental state and processes which are used to obtain the element.
- Predicting how ideas of green chemistry such as minimising the amount of unusable waste products, energy use and using more environmentally friendly chemical processes will affect the environment.



<u>CHEMnetBASE</u> provides a comprehensive database of chemicals and their uses, as well as chemical equations and properties.

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